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# An azobenzene-based photothermal energy storage system for co-harvesting photon energy and low-grade ambient heat *via* a photoinduced crystal-to-liquid transition

Liqi Dong<sup>1,2</sup> , Fei Zhai<sup>1</sup>, Hui Wang<sup>1,2</sup>, Cong Peng<sup>1</sup>, Yiyu Feng<sup>1</sup> , Wei Feng<sup>1,3,\*</sup>

<sup>1</sup>School of Materials Science and Engineering, Tianjin University, Tianjin 300072, China.

<sup>2</sup>Key Laboratory of Pollution Exposure and Health Intervention of Zhejiang Province, Interdisciplinary Research Academy, Zhejiang Shuren University, Hangzhou 310015, Zhejiang, China.

<sup>3</sup>Tianjin Key Laboratory of Composite and Functional Materials, Tianjin 300072, China.

\*Correspondence to: Prof./Dr. Wei Feng, School of Materials Science and Engineering, Tianjin University, No. 135, Yaguan Road, Haihe Education Park, Tianjin 300072, China. E-mail: weifeng@tju.edu.cn

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## Abstract

Ambient heat, slightly above or at room temperature, is a ubiquitous and inexhaustible energy source that has typically been ignored due to difficulties in its utilization. Recent evidence suggests that a class of azobenzene (Azo) photoswitches featuring a reversible photoinduced crystal-to-liquid transition could co-harvest photon energy and ambient heat. Thus, a new horizon has been opened for recovering and regenerating low-grade ambient heat. Here, a series of unilateral *para*-functionalized photoinduced liquefiable Azo derivatives is presented that can co-harvest and convert photon energy and ambient heat into chemical bond energy and latent heat in molecules and eventually release them in the form of high-temperature utilizable heat. A straightforward crystalline-to-liquid phase transition achieved with ultraviolet light irradiation (365 nm) is enabled by appending a halogen/alkoxy group on the *para*-position of the Azo photoswitches, and the release of thermal energy is triggered by short-wavelength visible-light irradiation (420 nm). The phase transition properties of the *trans*- and *cis*-isomers and the energy density, storage lifetime and heat release performance of the *cis*-liquid are investigated with differential scanning calorimetry, ultraviolet-visible absorption spectroscopy, and an infrared (IR) thermal camera. The experimental results indicate a high energy density of 335 J/g, a long lifetime of 5 d and a heat release of up to 6.3 °C due to the coupled photochemical-thermophysical mechanism. This work presents a new model for utilizing



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renewable energy, i.e., the photoinduced conversion of ambient thermal energy.

**Keywords:** Azobenzene, photoinduced crystal-to-liquid phase transition, photoisomerization, ambient heat conversion, light energy

## INTRODUCTION

Ambient heat, slightly above or at room temperature (RT), is a non-stop and omnipresent energy source that has typically been ignored due to its difficult utilization<sup>[1]</sup>. Ambient heat sources mainly include thermal radiation from the sun and low-grade waste heat generated by human activities, such as the waste heat from industrial production, the heat produced by refrigeration equipment, and the heat dissipation of electronic devices and vehicle exhausts. These massive amounts of ambient heat cannot yet be collected *via* economically viable methods, thus causing thermal pollution and resource waste<sup>[2-4]</sup>. Therefore, technologies that can effectively convert this ambient heat into useful energy are critically needed to achieve its potential in reducing energy consumption and mitigating environmental problems.

Various technologies have been proposed to harvest thermal energy, such as thermoelectrics<sup>[4,5]</sup>, thermochemical<sup>[6]</sup> and thermophysical<sup>[7]</sup> approaches, and heat pumps<sup>[8]</sup>. Phase change heat storage can convert relatively large amounts of ambient heat into latent heat over a very narrow temperature change, making it very suitable for recovering low-grade ambient heat<sup>[9-11]</sup>. However, phase change materials (PCMs) have an intrinsic limitation, in that the stored heat cannot be released in the form of higher temperatures. Moreover, conventional PCMs present difficulties in controlling the temperatures of heat absorption/release and the heat storage time. With the heat source removed, the ambient temperature drops below the melting or crystallization points and liquefied PCMs can immediately crystallize and lose the latent heat. These uncontrollable processes have constantly challenged long-term latent heat storage in conventional PCMs<sup>[12-14]</sup>. To date, a physical/chemical solution to gather and convert ambient heat to higher temperatures remains lacking. Thus, it is desirable to develop new technologies or materials to harvest, convert and store ambient heat and eventually release the stored thermal energy controllably.

Most recently, photoliquefiable azobenzene (Azo) molecules have shown significant potential for harvesting and storing thermal energy due to their reversible photoinduced crystal-to-liquid transitions (PCLTs), controllable heat storage and release, and zero gas/chemical emissions<sup>[15-19]</sup>. Han *et al.* discovered three heat storage-release schemes for storing thermal energy in liquid-state *cis*-isomers<sup>[20]</sup>. It is possible to preserve the latent heat for longer than two weeks in the *cis*-liquid isomers at temperatures below 0 °C unless triggered by visible light. In a groundbreaking photochemical-thermophysical system, Li and colleagues developed pyrazolylazophenyl ether molecules as solar and heat energy harvesters<sup>[21]</sup>. The energy capacities of these molecules are believed to be higher than that of conventional solar or heat energy storage methods based purely on phase transition or molecular photoisomerization. Furthermore, Xu *et al.* reported a novel photochromic dendrimer obtained by grafting Azo units onto dendrimers, which exhibited excellent solar energy storage, controlled heat release, self-repair and controllable adhesive switching properties<sup>[22,23]</sup>. Moreover, the research group of Han successfully designed new photoliquefiable Azo molecules featuring photocontrolled latent heat storage<sup>[24]</sup>. Therefore, Azo photoswitches with PCLT capacities have been investigated as promising energy harvest and storage materials. However, research on harvesting and storing ambient heat *via* PCLTs has not received significant recognition and a photoinduced conversion mechanism of ambient heat into higher temperatures has not been established.

Therefore, we report a series of unilateral *para*-functionalized photoliquefiable Azo derivatives capable of co-harvesting photon energy and ambient heat and generating high-grade thermal energy. These molecules constitute a solution to recycle and regenerate ambient heat *via* a photoinduced phase transition, which tactfully couples photochemistry and thermophysics. The unilateral functionalization drastically breaks the symmetry of Azo molecules, leading to a large melting point gap between the *trans* and *cis* configurations<sup>[20,25]</sup>. This molecular asymmetry contributes to *cis*-liquid phase stability upon cooling until the crystallization and release of heat triggered by specific light. The phase transition properties of *trans* and *cis* isomers and the energy density and storage lifetime of *cis*-liquid form were investigated with ultraviolet-visible (UV-Vis) absorption spectroscopy and differential scanning calorimetry (DSC). The optically-controlled phase transition is demonstrated by the selective crystallization of *cis*-liquid films. The heat release performance of the *cis*-liquid form is monitored using an infrared (IR) thermal camera. Finally, the mechanism by which photon energy triggers the conversion of ambient heat to higher temperatures within the photoliquefiable molecules is discussed. The design principles of the novel energy harvest and storage molecules are also discussed.

## EXPERIMENTAL

### Reagents and materials

In this study, nitrosobenzene (C<sub>6</sub>H<sub>5</sub>NO, 98%), 4-chloroaniline (C<sub>6</sub>H<sub>6</sub>ClN, 98%), 4-bromoaniline (C<sub>6</sub>H<sub>6</sub>BrN, 99%), 4-iodoaniline (C<sub>6</sub>H<sub>6</sub>IN, 98%), *para*-anisidine (C<sub>7</sub>H<sub>9</sub>NO, 99%), 4-ethoxyaniline (C<sub>8</sub>H<sub>11</sub>NO, 98%), glacial acetic acid (AR) and a microscope glass slide (24mm × 24mm × 0.17 mm, length × width × thickness) were purchased from Shanghai Aladdin Biochemical Technology Co., Ltd. Dichloromethane (DCM, AR and SP), ethyl acetate (AR) and petroleum ether (AR) were purchased from Tianjin Kemiou Chemical Reagent Co., Ltd. All reagents were used directly without further purification.

### Synthesis and characterization

Azo compounds were synthesized *via* the Mills reaction<sup>[26,27]</sup>. The detailed synthesis procedures and characterization data, including the high-resolution mass spectra (HRMS), <sup>1</sup>H and <sup>13</sup>C nuclear magnetic resonance (NMR) spectra, and Fourier transform-iIR spectra of the products, are included in [Supplementary Table 1](#).

A SepaBean machine automated flash chromatography system (Santai Technologies, Inc.) was used for purification. <sup>1</sup>H and <sup>13</sup>C NMR spectra were recorded using a 400 MHz spectrometer (INOVA, Varian, USA) with chloroform-*d* as the solvent and tetramethylsilane as an internal standard. IR spectra were recorded using a Bruker Vertex 70 Fourier spectrometer with the powdered samples pressed into KBr pellets. HRMS (ESI) were recorded on a Bruker solanX 70. The samples were charged/discharged under controlled light irradiation with different wavelengths, intensities, and durations. The intensity of light was measured with an optical power meter (CEL-NP2000, Beijing China Education Au-light Co., Ltd.). Optical microscope images were obtained using a Nikon SMZ745T (Japan). An IR thermal imaging camera (Fluke TiX640, USA) was used to monitor the heat released from the samples.

### UV-Vis absorption spectroscopy

The UV-Vis absorption spectra (1 cm path-length quartz cuvettes, DCM) were measured using a UV-Vis spectrophotometer (3600 Plus, Shimadzu, Japan). Baselines were corrected and spectra were normalized using Origin8.6 software.

For the solution state test, *trans*-isomers were dissolved in DCM. The UV-Vis spectra were then recorded as *trans*-isomers. After that, the samples were irradiated with 365-nm wavelength light (80 mW/cm<sup>2</sup>) in solution state until they reached a photostationary state (PSS), which generally takes 10 to 30 s. The UV-Vis

spectra were then recorded as the *cis*-isomers. Since the solution-state Azo compounds have ample free space, the *cis*-isomer content can exceed 95% in the 365-nm PSS. Therefore, we approximate the 365-nm PSS Azo compounds in the solution state as 100% *cis*-configurations in this work. The UV-Vis absorption spectra of *trans*- and *cis*-Azo compounds in solution state ( $1 \times 10^{-4}$  mol L<sup>-1</sup> DCM solution) can be seen in [Supplementary Figures 1, 2A, E and I](#).

For the solid-state test, a small sample was taken with a glass rod ( $\Phi = 0.5$  mm) and dissolved in a cuvette filled with DCM solution. The UV-Vis spectra were then recorded. The obtained data were normalized using Origin 8.6 software. The *cis*-isomer content in the 365-nm PSS can be estimated as the percentage change of absorbance at the wavelength of the  $\pi$ - $\pi^*$  transition peak following the method reported in the literature<sup>[28,29]</sup>.

$$\%cis \approx (A - A_{trans}) / (A_{cis} - A_{trans}) \quad (1)$$

The 365-nm PSS in dilute solutions was assumed to be ~100% *cis*, while the thermal-stationary state was assumed to be ~100% *trans*.  $A$  is the absorbance during the *trans*-to-*cis* photoisomerization.

### Energy charging process

In a dark room, *trans*-crystal powder samples were set on a 24 × 24 mm glass slide. The slide was set on a constant temperature heating platform that simulated the ambient heat ( $T_1$ ). The sample was then irradiated with 365-nm wavelength light (80 Mw/cm<sup>2</sup>, 5 cm away) until the *trans*-crystal was converted into the *cis*-liquid through photoisomerization. After that, the heating platform was adjusted to the lowest temperature ( $T_2$ ) so that the *cis* isomer could maintain a liquid state to maximize the *trans*-to-*cis* photoisomerization and obtain as much *cis*-isomer as possible. Images of the samples in the crystal and liquid states were recorded using a smartphone (Honor 30 Pro). During the irradiation, the UV-Vis absorption spectra of the samples were recorded at regular intervals. The obtained data were normalized by software. The  $T_1$  and  $T_2$  of the five Azo compounds are presented in [Supplementary Table 2](#).

### DSC measurements

DSC analysis was conducted on a DSC Q20 (TA Instruments, USA) with cooling from an RCS90. If not otherwise stated, the cooling and heating rates were set to 10 °C/min. In the phase transition property experiments, the *cis*-isomer samples were prepared using the solvent-assisted charging procedure to obtain the values of  $T_m$  and  $T_c$  as precisely as possible. The specific procedure is as follows. Each *trans*-isomer was dissolved in DCM and irradiated with 365-nm wavelength light (80 mW/cm<sup>2</sup>) until reaching the PSS. The solutions were then concentrated and dried in a vacuum before being transferred to DSC pans for analysis. In the energy density measurements experiments, the preparation procedure of the *cis*-isomer samples was the same as that for the energy charging process.

### Lifetime measurements

The *cis*-liquid samples of the 365-nm PSS were transferred to a dark room at a constant temperature (35 °C for Azo-Cl, 45 °C for Azo-Br, 65 °C for Azo-I and 25 °C for Azo-OMe and Azo-OEt). The UV-Vis absorption spectra of the samples were recorded at regular intervals until the samples reverted to their initial state, *i.e.*, the *trans*-crystal state. The obtained data were normalized by software.

### Photoinduced liquid-to-crystal phase transition experiments

*Cis*-liquid isomer thin films were prepared by placing 8 mg of *trans*-crystal isomer sample on a clean glass slide (24 × 24 mm) and melting it on a heating platform. The *trans*-liquid sample was flattened by pressing with another glass slide to fill the entire space between them and cooled to  $T_1$  [Supplementary Table 2](#) to

simulate the ambient heat. The sample was then irradiated with 365-nm wavelength light ( $80 \text{ mW/cm}^2$ ) to induce the *trans*-to-*cis* photoisomerization. The temperature was then reduced to  $T_2$  [Supplementary Table 2](#) using an ethanol bath with a circulatory system and the light irradiation was removed. Half of the sample was covered with cardboard and irradiation was applied with 420-nm wavelength light ( $80 \text{ mW/cm}^2$ ) to trigger the *cis*-to-*trans* isomerization and induce crystallization. After crystallization, the cover was removed and images were taken with an optical microscope at room temperature. The UV-Vis absorption spectra of the samples were recorded at regular intervals during the photoinduced crystallization. The obtained data were normalized by software.

### Ir thermal imaging

The *cis*-liquid samples were set in a dark room at a constant temperature ( $35 \text{ }^\circ\text{C}$  for Azo-Cl,  $45 \text{ }^\circ\text{C}$  for Azo-Br,  $65 \text{ }^\circ\text{C}$  for Azo-I and  $25 \text{ }^\circ\text{C}$  for Azo-OMe and Azo-OEt). The samples were then irradiated with 420-nm wavelength light ( $80 \text{ mW/cm}^2$ , 5 cm away) under an IR thermal imaging camera.

### Theoretical calculations

The structures of molecules 1 to 5 were first optimized using the density functional theory (DFT) at the  $\omega$  B97XD/def2-SVP level. Among all the possible conformations for both the *trans*- and *cis*-isomers, all geometry optimizations were performed using the Gaussian 16a package. Harmonic frequency calculations were performed using the same theory to help verify that none of the structures have imaginary frequencies. The most stable structures were selected to analyze the direction and magnitude of the molecular dipole moment.

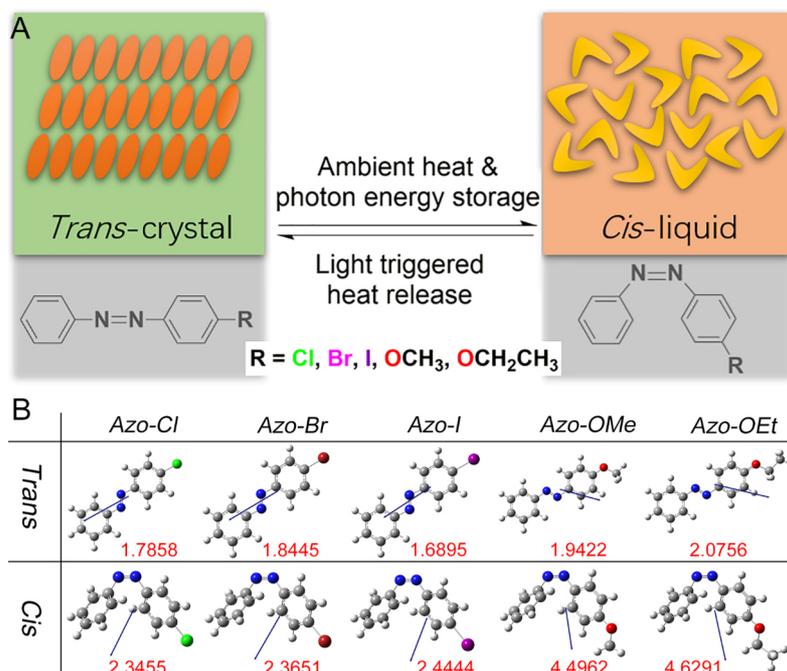
## RESULTS AND DISCUSSION

### Molecular design

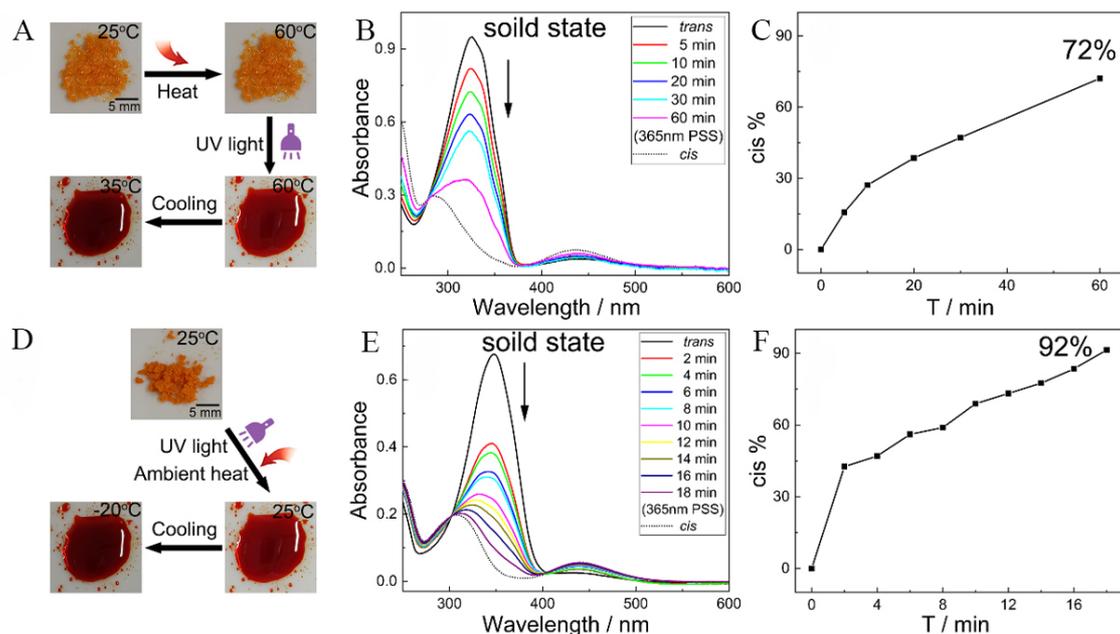
In order to achieve PCLTs, the Azo molecules must have a large melting point gap ( $\Delta T_m$ ) between the *trans*- and *cis*-isomers so that the *trans*-crystal and *cis*-liquid phases can be formed under the specific temperature  $T_{phase}$  ( $cis-T_m < T_{phase} < trans-T_m$ ). Grafting functional groups on the benzene ring of the Azo molecules disrupts the molecule symmetry, resulting in a large  $\Delta T_m$  between two isomers. In addition, the DFT calculation result shows that the *cis* isomers of the unilateral-functionalized Azo derivatives have significantly higher polarity than their *trans*-forms [[Figure 1B](#)]. The high polarity of the *cis*-isomers effectively disrupts the  $\pi$ - $\pi$  interactions between the aromatic groups, eventually reducing their ability to pack and form ordered crystals.<sup>[20]</sup> Hence, we synthesized five Azo photoswitches that crystallize in the *trans*-form and liquefy in the *cis*-form at or slightly above RT for co-harvesting photon energy and ambient heat via PCLTs [[Figure 1A](#)].

### Energy charging: photoinduced *trans*-crystal-to-*cis*-liquid transition

The photoliquidification behaviors of all five Azos forms were observed and presented in [Figure 2](#) and [Supplementary Figure 2](#). To absorb ambient heat above RT, we selected halogen-grafted Azos with photoinduced phase transition temperatures above RT. Such Azos liquify under UV light at ambient temperatures above RT and store the ambient heat in liquid-state molecules as latent heat [[Figure 2A](#), [Supplementary Figure 2D and H](#)]. Taking the Azo-Cl molecules as examples, as shown in [Figure 2A](#), an ambient temperature of  $60 \text{ }^\circ\text{C}$  is first simulated. The *trans*-crystal sample is then irradiated with 365-nm wavelength light to trigger the *trans*-to-*cis* photoisomerization and the simultaneous crystal-to-liquid phase transition, which harvests the photon energy and ambient heat. During light irradiation, Azo-Cl was sampled and dissolved, and its isomerization degree was monitored by UV-Vis spectroscopy.



**Figure 1.** (A) Concept and molecular design for co-harvesting of ambient heat and photon energy via the *trans*-crystal-to-*cis*-liquid transition of photoliquefiable Azos. (B) Dipole moment calculation (Debye) of photoliquefiable Azos.



**Figure 2.** Energy harvested from ambient heat and photons, i.e., the energy charging process. First column: images of Azo-Cl (A) and Azo-OEt (D) samples on glass slides, showing the *trans*-crystals at the specific temperature, *cis*-liquid after UV light irradiation (365 nm) and *cis*-liquid after cooling, as indicated by the arrows. Second column: time-evolved UV-Vis absorption spectra of Azo-Cl (B) and Azo-OEt (E) under UV light of 365 nm. The spectra were normalized with respect to the isosbestic point at 279 nm for Azo-Cl and 302 nm for Azo-OEt, respectively. The arrows indicate the order of tests performed. Third column: the *cis*-isomer percentage of Azo-Cl (C) and Azo-OEt (F) vs. irradiation time.

For the purpose of maximizing the *trans*-to-*cis* photoisomerization and obtaining as much *cis*-isomer as

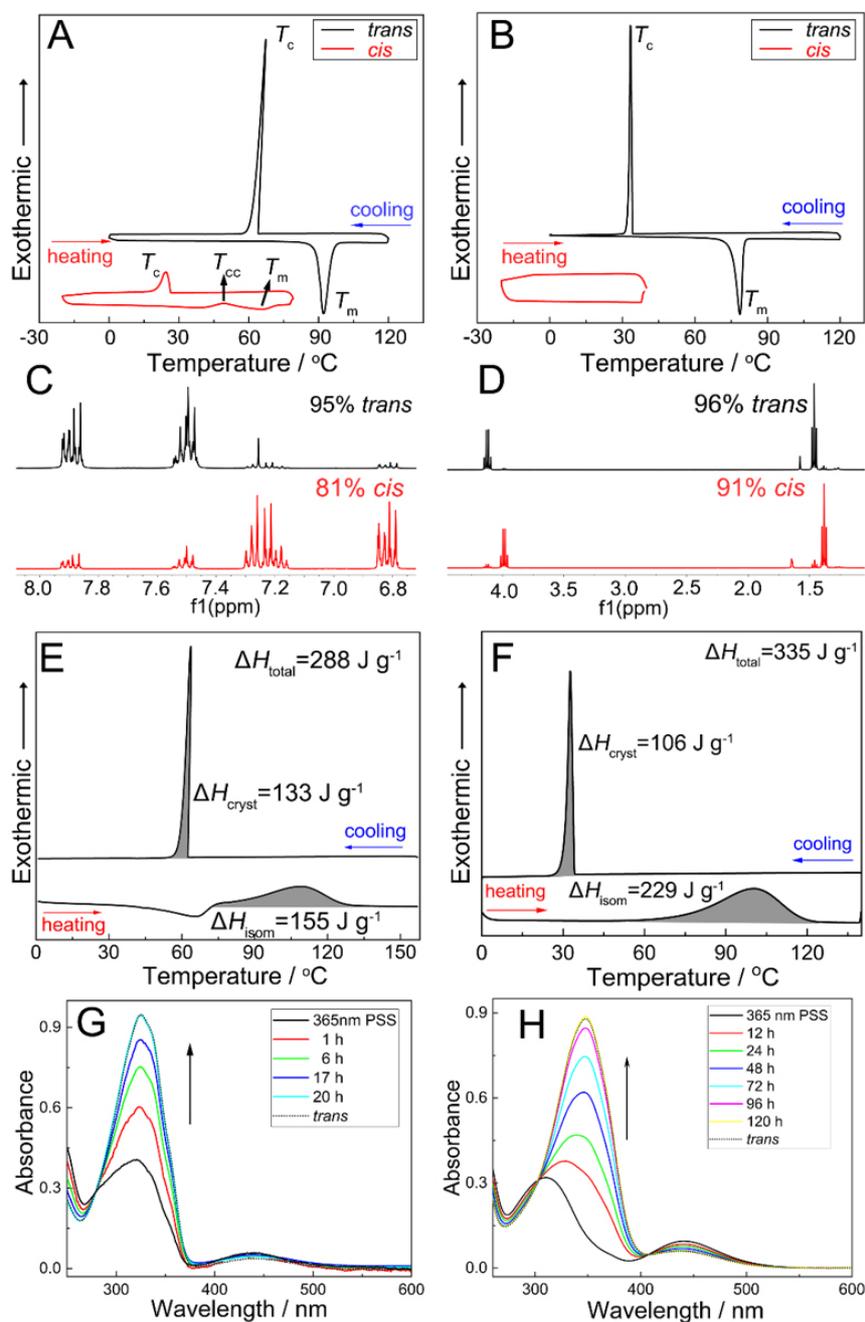
possible, the *cis*-liquid sample is cooled to the lowest temperature ( $T_2$  in [Supplementary Table 2](#)) at which it could remain in the liquid state while maintaining light irradiation. As shown in [Figure 2B](#) and [C](#), the Azo-Cl sample reached the 365-nm PSS after a 60 min irradiation and the *cis*-isomer proportion reached 72%. It should be noted that prolonging the irradiation could reduce the *cis*-liquid content. The main reason for this is that the relatively high temperature of ambient heat causes the reverse conversion of *trans*-to-*cis* isomerization [[Supplementary Figure 2B, C, F and G](#)]. For ambient heat at RT, we selected alkoxy-grafted Azos with photoinduced phase transition temperatures at RT. Taking the Azo-OEt molecules as examples, as displayed in [Figure 2D](#), the *trans*-isomer of Azo-OEt is photoliquidified into a reddish-orange liquid under RT. The Azo-OEt sample could reach the 365-nm PSS within 18 min at RT and exhibit a high *trans*-to-*cis* photoisomerization degree with a *cis*-isomer content of 92% [[Figure 2E and F](#)]. This is approaching the maximum level achievable in diluted solutions (91%-96%).<sup>[25]</sup> Moreover, The 365-nm PSS irradiation time and *cis*-isomer content of Azo-OMe sample are 10 min and 86% [[Supplementary Figure 2J and K](#)]. Through such high effective *trans*-to-*cis* isomerization during photo liquefaction, energy harvesting could be more efficient and could result in high energy densities. Furthermore, the alkoxy-grafted *cis*-isomers remained in their liquid phase at -20 °C, indicating that the phase state of the *cis*-isomers is highly reliant on the *para*-substituent [[Figure 2D and 2L](#)]<sup>[30,31]</sup>.

### Energy storage

The notable difference in crystallinity between the *trans*- and *cis*-isomers of the five as-prepared Azos, *i.e.*, Azo-Cl [[Figure 3A](#)], Azo-Br [[Supplementary Figure 3A](#)], Azo-I [[Supplementary Figure 3B](#)], Azo-OMe [[Supplementary Figure 3C](#)] and Azo-OEt [[Figure 3B](#)], was measured by DSC. The *trans*-forms of all five Azos exhibit sharpened melting peaks upon heating and crystallization peaks upon cooling. The same melting and crystallization characteristics are significantly reduced (Azo-Cl, Azo-Br, and Azo-I molecules) or absent (Azo-OMe and Azo-OEt compounds) for the *cis*-form counterparts. This intriguing result could be attributed to the significantly higher polarity of the *cis*-isomers than the *trans*-isomers, which effectively disrupts the  $\pi$ - $\pi$  interactions between aromatic groups, eventually reducing their ability to pack and form ordered crystals.<sup>[20,32]</sup> The *cis*-forms of halogen-grafted Azo-Cl [[Figure 3A](#)], Azo-Br [[Supplementary Figure 3A](#)] and Azo-I [[Supplementary Figure 3B](#)] maintained their liquid phase at temperatures above 24, 46, and 64 °C, respectively. These temperatures were slightly lower than the observed ones for maintaining the *cis*-liquid phase. This discrepancy may be attributed to the different preparation methods (solvent-assisted or direct irradiation in the solid state), resulting in different *cis*-isomer contents [[Supplementary Figures 3C-E](#)]. Additionally, Azo-Cl exhibits a cold-crystallization peak at ~45 °C, indicating that it is more prone to crystallization than other compounds [[Figure 3A](#)].

The *cis*-forms of the alkoxy-grafted Azos displayed remarkable liquid-phase stability at temperatures below 0 °C, thereby providing opportunities for low-temperature heat release [[Figures 3B and Supplementary Figure 3C](#)]. As summarized in [Supplementary Table 3](#), these data match the observed phenomena, suggesting that direct irradiation in the solid state has the same effect on energy storage as the solvent-assisted charging method at RT [[Figures 3D and Supplementary Figure 3F](#)]. Therefore, a significant difference can be observed between the *trans*- and *cis*-isomers for all Azos. Specifically, the *trans*-isomers are highly crystalline, while the *cis*-isomers are liquid at the same temperature. This phase state discrepancy indicates an immense opportunity for co-storing the photon energy and ambient heat in the *cis*-liquid form of molecules and releasing them as high-grade thermal energy through the *cis*-liquid-to-*trans*-crystal transition.

DSC was also used to investigate the amount of energy stored in the *cis*-liquid isomers of all five Azos: Azo-Cl [[Figure 3E](#)], Azo-Br [[Supplementary Figure 4A](#)], Azo-I [[Supplementary Figure 4B](#)], Azo-OMe [[Supplementary Figure 4C](#)] and Azo-OEt [[Figure 3F](#)]. Each Azo compound has two separate exothermic



**Figure 3.** Phase transition of isomers and energy storage properties. First row: DSC plots of *trans* (top black curve)- and *cis* (bottom red curve)-isomers upon heating and cooling with Azo-Cl (A) as a representative example of halogen-grafted Azos and Azo-OEt (B) as a representative example of alkoxy-grafted Azos.  $T_m$ ,  $T_c$ , and  $T_{cc}$  represent the melting, crystallization, and cold-crystallization temperatures, respectively. Second row:  $^1\text{H}$  NMR spectra of Azo-Cl (C) and Azo-OEt (D) isomers for DSC test. Third row: DSC curves of typical *cis*-liquid Azos for Azo-Cl (E) as a representative example of halogen-grafted Azos and Azo-OEt (F) as a representative example of alkoxy-grafted Azos. Fourth row: time-evolved UV-Vis absorption spectra of the *cis*-liquid form of Azo-Cl (G) and Azo-OEt (H) reversion in darkness (the spectra were normalized with respect to their corresponding isosbestic point). The arrows indicate the test order.

peaks. The broad exothermic heat flows in the first heating segment were attributed to the heat-induced *cis*-to-*trans* isomerization. This portion of energy came from the photon energy captured during the energy charging process. The sharp exothermic peaks in the cooling segment were ascribed to the crystallization of

the *trans*-liquid state. This portion of energy came from the thermal energy absorbed from the ambient environment during the energy charging process. Therefore, these Azo molecules harvested and stored the hard-to-collect ambient heat through their PCLT behaviors without notably compromising the storage capacity of photon energy. The photon energy and ambient heat were tactfully integrated into the single-component small molecule *via* photochemistry and thermophysics, resulting in a high energy density (172–335 J/g). The energy densities of all five Azo compounds are presented in Table 1. Furthermore, the time-evolved UV-Vis absorption spectra of *cis*-liquid Azos in darkness were plotted to survey their storage lifetimes (Azo-Cl in Figure 3G, Azo-Br in Supplementary Figure 4D, Azo-I in Supplementary Figure 4E, Azo-OMe in Supplementary Figure 4F and Azo-OEt in Figure 3H). The results are summarized in Table 1. Alkoxy-grafted Azos have a significantly longer storage lifetime than halogen-grafted Azos because the *cis*-form halogen-grafted Azos require higher ambient temperatures to sustain their liquid phase. However, the high temperature could also cause thermally-induced *cis*-to-*trans* isomerization. Therefore, it is necessary to select suitable photoliquefiable Azos to accommodate various application scenarios. For example, halogen-grafted Azos are potentially suitable for applications that need to accumulate heat in a short period of time or within a small area, while alkoxy-grafted Azos are more favorable candidates for long-term energy storage.

### Energy discharging: photoinduced *cis*-liquid-to-*trans*-crystal transition

Further verification of the photoinduced crystallization is achieved *via* selectively exposing the *cis*-liquid films to visible light (420 nm) using an optical mask (typically cardboard), as shown in Figure 4A. The 420-nm visible light irradiation can help the *cis*-isomers overcome the energy barrier and accelerate the *cis*-liquid-to-*trans*-crystal isomerization. As shown in Figure 4B, C and Supplementary Figure 5A–C, the *cis*-liquid films are partially covered with cardboard on the left and the entire films are exposed to light irradiation at 420 nm. The phase of the covered part (the left side) has no variation, whereas the right sides of the *cis*-liquid films crystallize due to light irradiation. These optical microscope images clearly visualize the crystalline features of each *trans*-isomer and the boundary between the original liquid phase and the generated crystalline phase. This indicates that the crystallization is induced exclusively by light irradiation and not artifact nucleation. Therefore, temporal and spatial control over the heat release can be achieved with light triggers. Supplementary Figure 6 shows the decrease in *cis* content in the liquid films under the 420-nm wavelength light irradiation. After 20 s of the 420-nm light irradiation, the *cis* content of the halogen-grafted Azos drops below 40% and crystallizes [Figure 4D, Supplementary Figure 5D and 5E]. In contrast, the alkoxy-grafted Azos require at least 10 min to achieve the same effects [Figure 4E and Supplementary Figure 5F], indicating a higher liquid phase stability.

For investigating the heat release performance, 420-nm light was used to irradiate the liquid films of the *cis*-isomers and the temperature changes were observed using an IR thermal imaging camera in real time. As indicated in Figure 4F, G and Supplementary Figure 5G–I, the samples became a source of heat in comparison to their surroundings while under light irradiation. The liquid films exhibited a rapid temperature increase until reaching the highest temperature, indicating that the stored photon energy and ambient heat were completely released. Afterwards, the sample temperature slowly reduced *via* dissipating heat into the surrounding environment [Figure 4H, I and Supplementary Figure 5J–L]. The temperature changes of the as-obtained Azos are presented in Table 1. Based on these results, it is clear that photoliquefiable Azos have excellent energy-storage capabilities and controlled on-demand heat release, thereby potentially contributing to ambient heat and solar energy utilization.

### Mechanism of photoinduced transition between *trans*-crystal and *cis*-liquid

This new energy conversion/storage principle takes advantage of the different melting points of Azo molecules with different configurations, which is a typical PCLT. Figure 5 presents the PCLT mechanism.

**Table 1. Energy density, storage lifetime and temperature changes of cis-form Azo compounds**

Azo compounds	Energy density				$\Delta H_{total}$ (J/g or kJ/mol)	Storage lifetime	Temper-ature changes (°C)
	Isomerization part		Crystallization part				
	$T_{isom}$ (°C)	$\Delta H_{isom}$ (J/g or kJ/mol)	$T_c$ (°C)	$\Delta H_{cryst}$ (J/g or kJ/mol)			
Azo-Cl	108	155 or 34	67	133 or 29	288 or 63	20 h	+6
Azo-Br	107	92or 24	75	97 or 25	189 or 49	18 h	+5.1
Azo-I	97	102 or 31	92	70 or 22	172 or 53	1 h	+4.0
Azo-OMe	96	155 or 33	32	76 or 18	231 or 51	5d	+4.7
Azo-OEt	100	229 or 52	33	106 or 24	335 or 76	4 d	+6.3

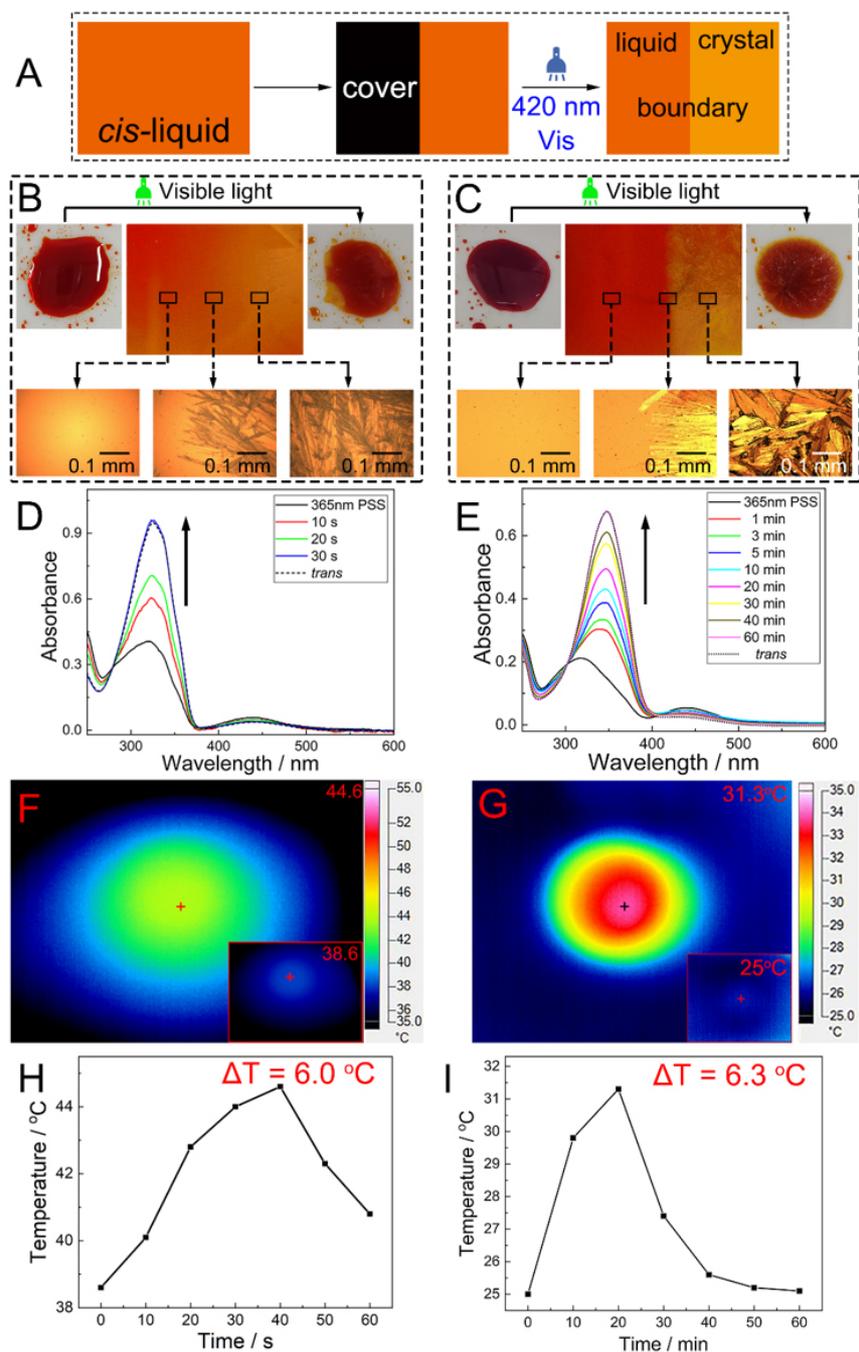
Generally, the melting point of *trans*-Azo molecules ( $trans-T_m$ ) is higher than that of the *cis*-Azo molecules ( $cis-T_m$ ). Assuming that the Azo molecules irradiated with UV light are at a temperature ( $T_{phase}$ ) between  $trans-T_m$  and  $cis-T_m$ , the *trans*-Azo would absorb the photon energy for photoisomerization from *trans*-crystal to *cis*-crystal and consume ambient heat spontaneously for phase transition from *cis*-crystal to *cis*-liquid. The ambient thermal energy can come from the waste heat from industrial production, the heat produced by refrigeration equipment, and the heat dissipation of electronic devices and vehicle exhausts. Consequently, photon energy and ambient heat are co-stored as N=N bond energy and phase change latent heat, respectively. If needed, the stored energy can be released as high-grade heat energy via *cis*-liquid-to-*trans*-liquid isomerization ( $\Delta H_{isom}$ ) and *trans*-liquid-to-*trans*-crystal crystallization ( $\Delta H_{cryst}$ ). Eventually, a system is formed with ambient heat and photon energy as inputs and high-grade thermal energy as output.

#### Photoinduced ambient heat conversion

We have proved that photoliquefiable Azo molecules can co-harvest and convert photon energy and ambient heat into chemical bond energy and latent heat in molecules, eventually releasing them as utilizable high-grade heat. This new model of renewable energy utilization is illustrated in Figure 6.

From the perspective of thermophysics, conventional PCMs are temperature-dependent and cannot convert ambient heat to a higher temperature. In contrast, photoliquefiable Azo molecules are light-dependent, which can convert ambient heat between two different phase transition temperatures (the high temperature  $trans-T_m$  and the low temperature  $cis-T_m$ ). In the *trans*-crystal-to-*cis*-liquid conversion, the phase transition temperature decreased from  $trans-T_m$  to  $cis-T_m$  alongside the *trans*-to-*cis* photoisomerization. Thus, the crystal-to-liquid phase transition occurred at a temperature below  $trans-T_m$ , allowing low-grade ambient heat to be harvested. In the *cis*-liquid-to-*trans*-crystal conversion, the *cis*-to-*trans* isomerization raised the phase transition temperature back (from  $cis-T_m$  to  $trans-T_m$ ), which enabled the crystallization to occur at a temperature above  $cis-T_m$ , thus releasing high-grade heat. In short, Azo molecules can liquefy and store thermal energy at  $cis-T_m$  but crystallize and release it at  $trans-T_m$ . Because  $trans-T_m$  is higher than  $cis-T_m$ , the heat output can exceed the input. Under such circumstances, designing Azo molecules with a large  $T_m$  difference between the *trans* and *cis* forms may be the key to such photothermal energy storage systems.

From the perspective of photochemistry, the essence of this hybrid energy system is the photoinduced ambient heat conversion. Whether it is the charging or discharging process, photon energy is indispensable to converting the difficult-to-use ambient heat into utilizable high-grade heat. During the charging process, the additional photon energy induced configuration change of the Azo molecules, thus harvesting or extracting the low-grade heat from the environment and storing them together. During the discharging process, the photon energy enabled the release of the stored energy as high-grade heat. Compared with



**Figure 4.** Experimental demonstration of selective crystallization of liquid sample and the heat release monitoring *via* photoinduction. (A) Schematic illustration of the selective crystallization of liquid sample *via* photoinduction. Selective photoinduced crystallization process of (B) Azo-Cl as a representative example of halogen-grafted Azos and (C) Azo-OEt as a representative example of alkoxy-grafted Azos, which were irradiated with visible light (420 nm) at ambient temperatures of 35 and -20 °C, respectively. The left side of the liquid film is covered with cardboard to preserve the *cis*-isomers in the stable liquid state. Optical microscope images of the left side, right side, and center of the film, indicating *cis*-isomers at the liquid phase, solid phase, and the liquid and crystalline solid boundary, respectively. The images on the top corners of (B) and (C) depict the *cis*-liquid samples (left) before and (right) after the 420-nm light irradiation. Normalized UV-Vis spectra of the uncovered parts for (D) Azo-Cl and (E) Azo-OEt during the 420-nm light irradiation. The arrows indicate the test order. IR thermal images of (F) Azo-Cl and (G) Azo-OEt under the 420-nm light irradiation. The insets show the samples before the light irradiation. The number in the upper right corner of each figure indicates the highest temperature of the materials. Temperature-time curves of (H) Azo-Cl and (I) Azo-OEt during heat release.  $\Delta T$  represents the temperature difference before and after heat release.

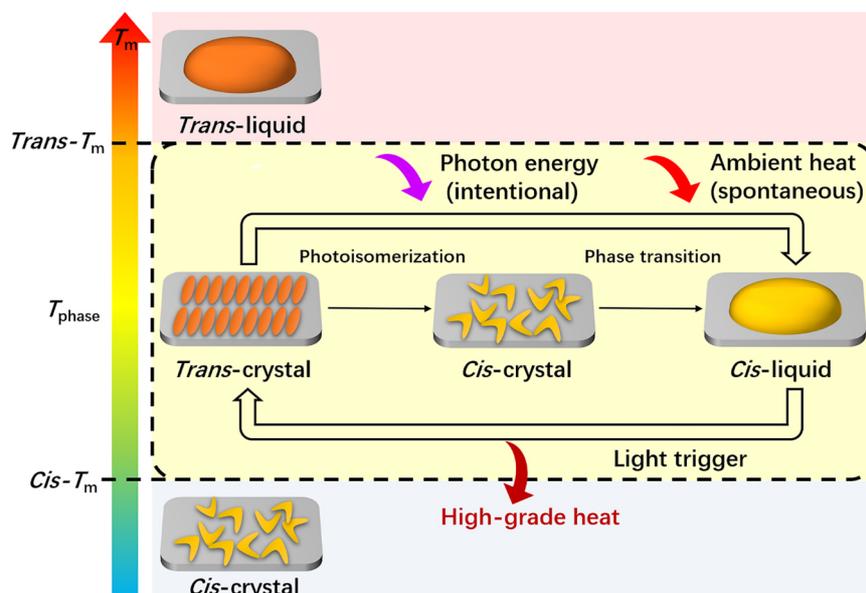


Figure 5. PCLT mechanism of photoliquefiable Azo molecules.

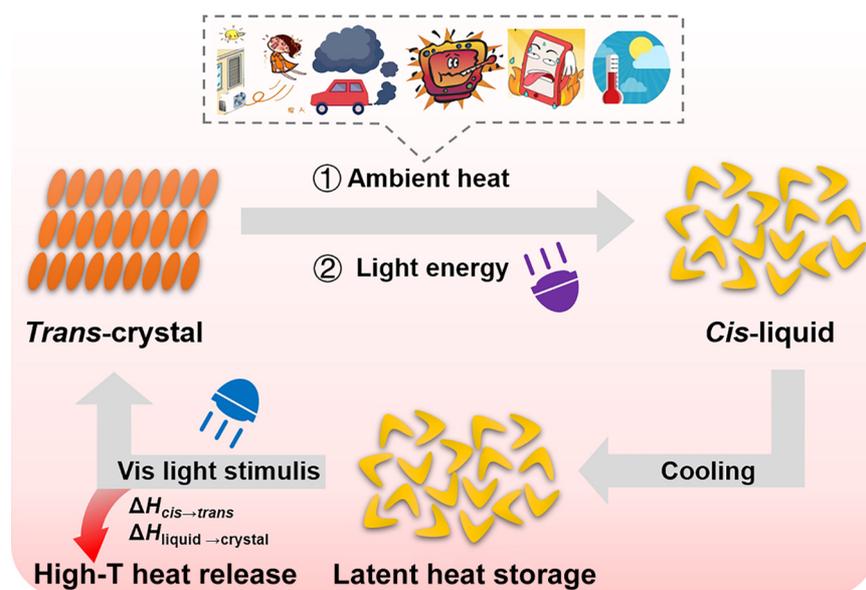


Figure 6. Schematic illustration of photoinduced ambient heat conversion model.

conventional heat pumps that convert ambient heat by spending high-grade energy (mechanical work or electricity), such photothermal energy storage systems are especially advantageous in terms of sustainability, environmental friendliness, and economics.

## CONCLUSIONS

In summary, we have developed five photoliquefiable Azo molecules in two categories, namely, halogen-grafted Azos for absorbing ambient heat above RT and alkoxy-grafted Azos for harvesting ambient heat at RT. These Azos can co-harvest photon energy and ambient heat and store them as chemical bond energy

and latent heat before releasing them as high-grade heat when needed. This novel photothermal energy harvest and storage system tactfully coupled photochemistry and thermophysics by exploiting the reversible PCLT feature of Azo molecules and achieved photoinduced ambient heat conversion. Unlike traditional molecular solar thermal systems, such a photochemical-thermophysical-coupled mechanism can offer an energy density up to 335 J/g, which is beyond that of systems based purely on phase transition or molecular photoisomerization. The photoinduced selective crystallization of *cis*-liquid isomers indicated that the heat release was controllable temporally and spatially, which was of great significance for applications requiring heat release on demand. Overall, this work provides a new tactic for recovering and regenerating low-grade ambient heat and paved the way for developing high-energy-density molecular solar thermal systems powered by co-harvesting photon energy and ambient heat. Future investigations may focus on the sunlight-induced *trans*-crystal-to-*cis*-liquid transition and photoinduced heat release at temperatures below 0 °C.

## DECLARATIONS

### Authors' contributions

conception and design of the study and wrote the manuscript: Dong L

Materials support: Dong L, Wang H, Peng C

Investigation, data analysis and interpretation: Dong L, Zhai F, Feng Y

Reviewed the manuscript, funding acquisition: Feng W

### Availability of data and materials

Not applicable.

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### Conflicts of interest

All authors declared that there are no conflicts of interest.

### Ethical approval and consent to participate

Not applicable.

### Consent for publication

Not applicable.

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