Supplementary Materials

Triethanolamine assisted synthesis of bimetallic nickel cobalt nitride/nitrogendoped carbon hollow nanoflowers for supercapacitor

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Figure S1. XRD patterns of N_0C_3 -TEOA, N_1C_2 -TEOA, N_1C_1 -TEOA, N_2C_1 -TEOA, N_3C_0 -TEOA.



Figure S2. SEM images of (A, B) N_0C_3 -TEOA, (C, D) N_1C_2 -TEOA, (E, F) N_1C_1 -TEOA, (G, H) N_2C_1 -TEOA, (I, J) N_3C_0 -TEOA.



Figure S3. N₂ adsorption-desorption isotherms of (A) N₀C₃-TEOA, (B) N₁C₁-TEOA,
(C) N₂C₁-TEOA, (D) N₃C₀-TEOA (Inset: Corresponding pore diameter distribution patterns).



Figure S4. FTIR spectrum of Ni₃N-Co₂N_{0.67}/NC, N₀C₃-TEOA, N₁C₂-TEOA, N₁C₁-TEOA, N₂C₁-TEOA and N₃C₀-TEOA.



Figure S5. XPS full spectrum of Ni₃N-Co₂N_{0.67}/NC.



Figure S6. Capacitive- and diffusive-controlled contributions of the Ni_3N -Co₂N_{0.67}/NC electrode by CV curves at (A) 10 mv s⁻¹, (B) 20 mV s⁻¹, (C) 30 mV s⁻¹ (D) 40 mV s⁻¹, (E) 50 mV s⁻¹.



Figure S7. (A) CV curves of active carbon at different sweep rates. (B) GCD profiles of active carbon at different current densities.

Table S1. Comparison of specific capacitance of Ni_3N -Co₂ $N_{0.67}/NC$ with other similar materials reported earlier.

Materials	Electrolyte	Potential window (V)	Current Density	Specific capacitance (F g ⁻¹)	Refs.
Ni ₃ N-Co ₂ N _{0.67} /NC	1 M KOH	0.5	1 A g ⁻¹	1582	This work
Ni-Co-N/GP	3 М КОН	0.5	4 A g ⁻¹	960	[1]
Ni-CoN@NC	1 M KOH	0.6	1 A g ⁻¹	993	[2]
Ni-doped Co-Co ₂ N	1 M KOH	0.5	2 mA/cm^2	181	[3]
B-Co ₄ N-20/NF	ЗМ КОН	0.5	1 A g ⁻¹	489	[4]
Co ₄ N@LOC/CC	1 M KOH	0.45	1 mA/cm^2	613.4	[5]
Co ₃ N	6 M KOH	0.43	$0.5 \mathrm{A g^{-1}}$	112.3	[6]
Co ₃ O ₄ @Ni-Co LDH	2 M KOH	0.55	$0.5 \mathrm{A g^{-1}}$	1318	[7]
NCH/EMCTs-36	6 M KOH	0.4	$0.5 \mathrm{A g^{-1}}$	801	[8]
Co(OH) ₂ /Ni	2 M KOH	0.4	1 A g ⁻¹	1310	[9]
Ni-Co LDH	1 M KOH	0.5	1 A g ⁻¹	1530	[10]
Ni/Co-MOF/rGO	1 M KOH	0.4	1 A g ⁻¹	1162	[11]
3D Ni–Co LDH/NiNw	6 M KOH	0.55	0.125 A g ⁻¹	466.6	[12]

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