

## Supplementary Materials

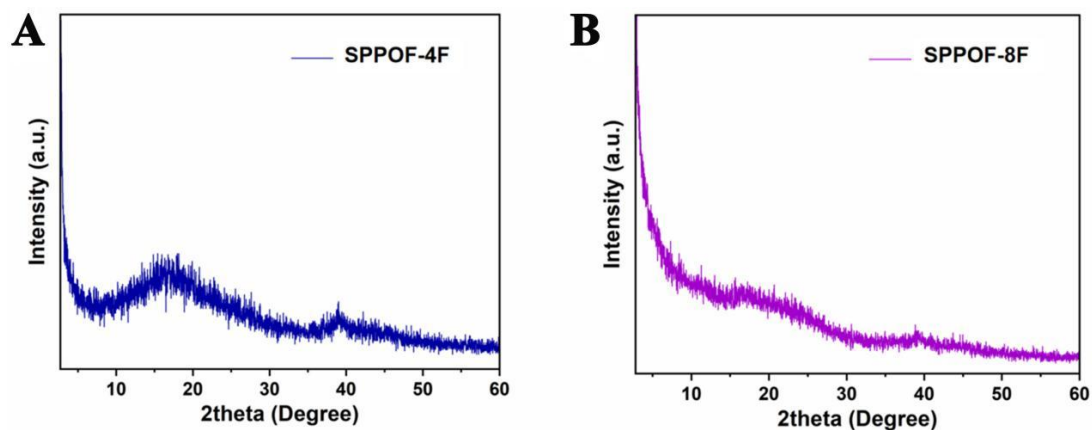
### Fluorinated porous organic frameworks for C<sub>2</sub>F<sub>6</sub>/CF<sub>4</sub> gases separation

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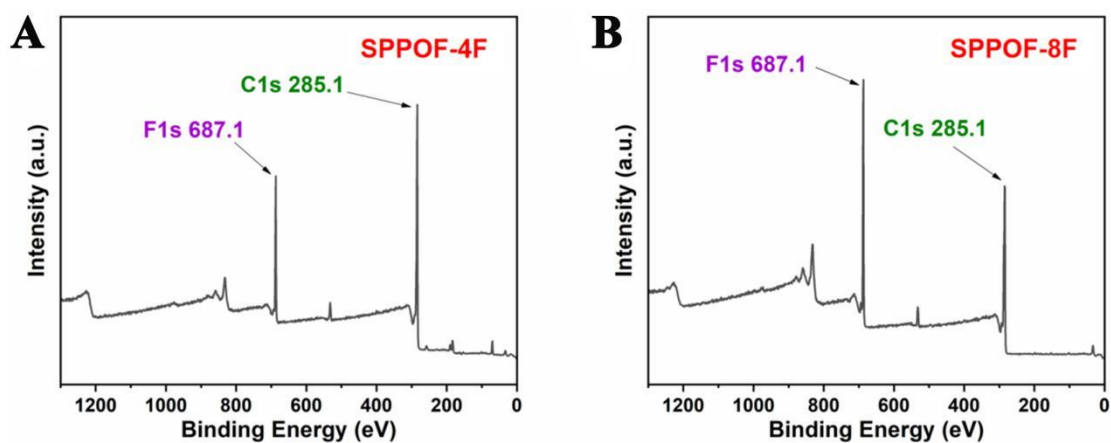
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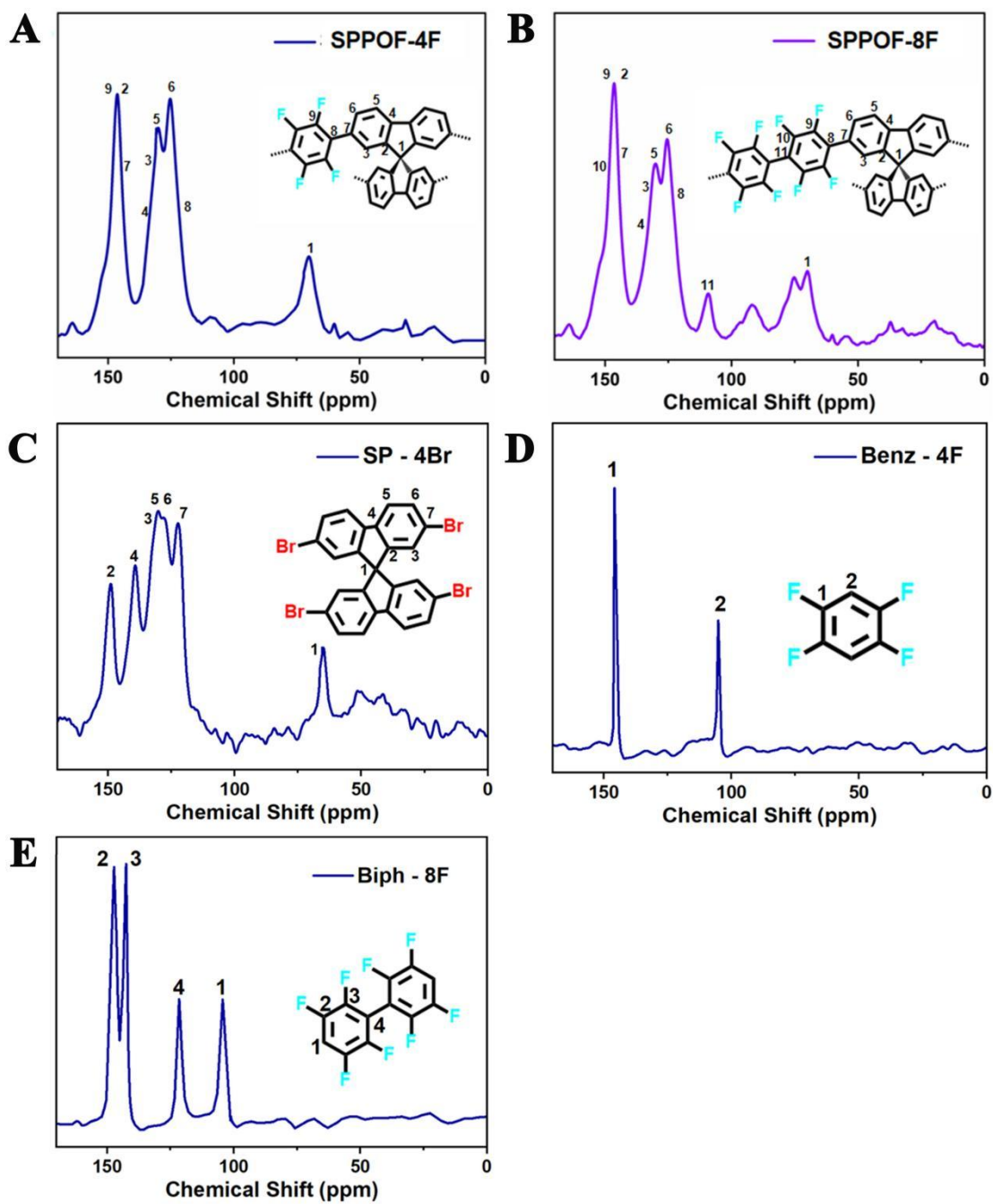
## Supplementary Figures



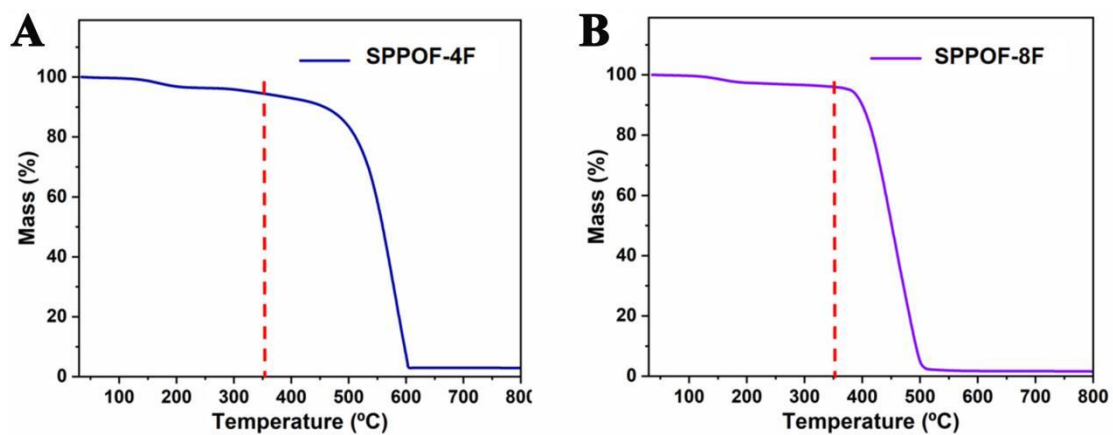
**Supplementary Figure 1.** Powder X-ray diffraction spectra of (A) SPPOF-4F and (B) SPPOF-8F.



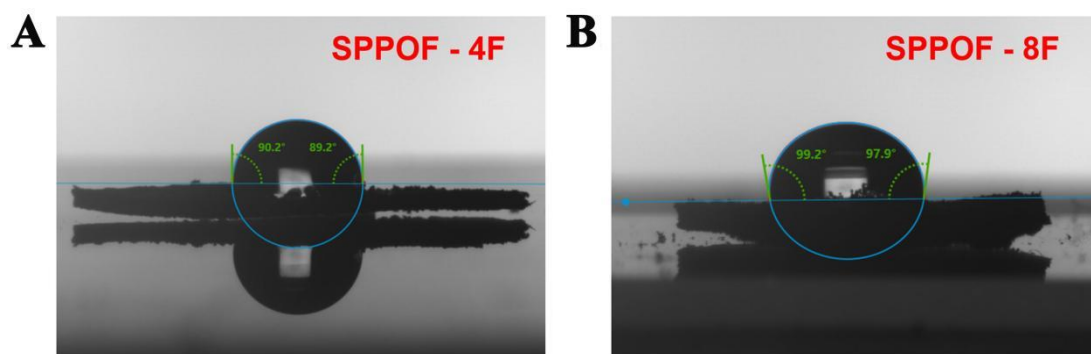
**Supplementary Figure 2.** XPS spectra of (A) SPPOF-4F and (B) SPPOF-8F.



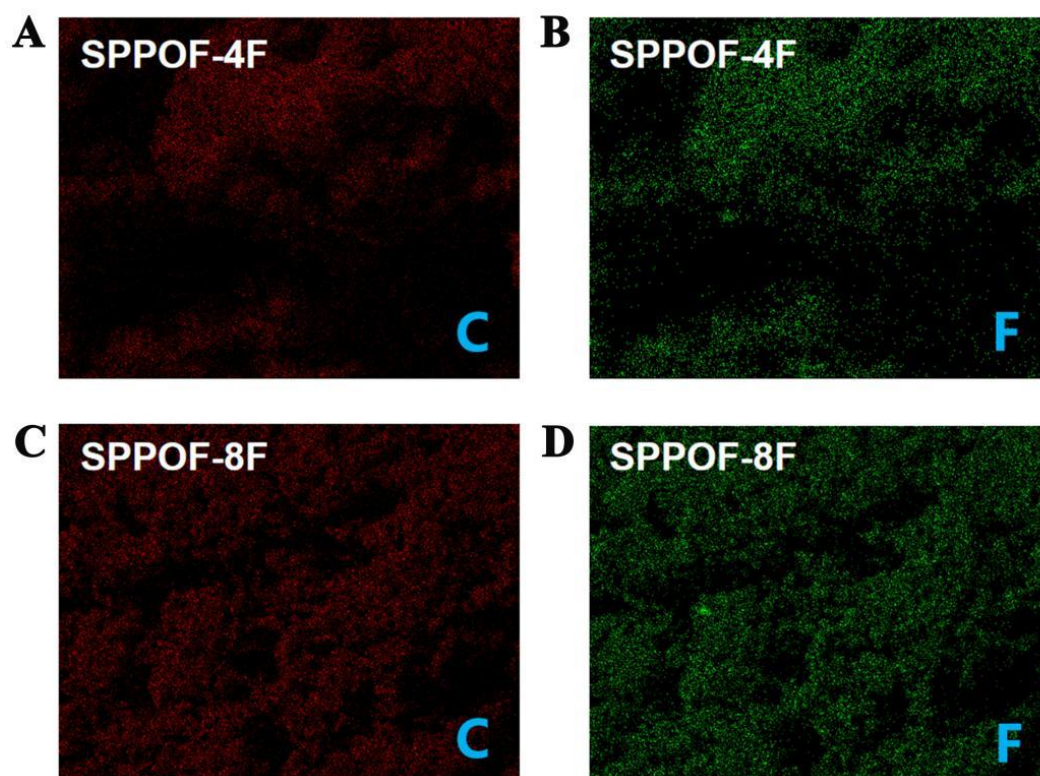
**Supplementary Figure 3.** Solid-state  $^{13}\text{C}$  CP/MAS NMR spectra: (A) SPPOF-4F, (B) SPPOF-8F, (C) SP-4Br, (D) Benz-4F and (E) Biph-8F.



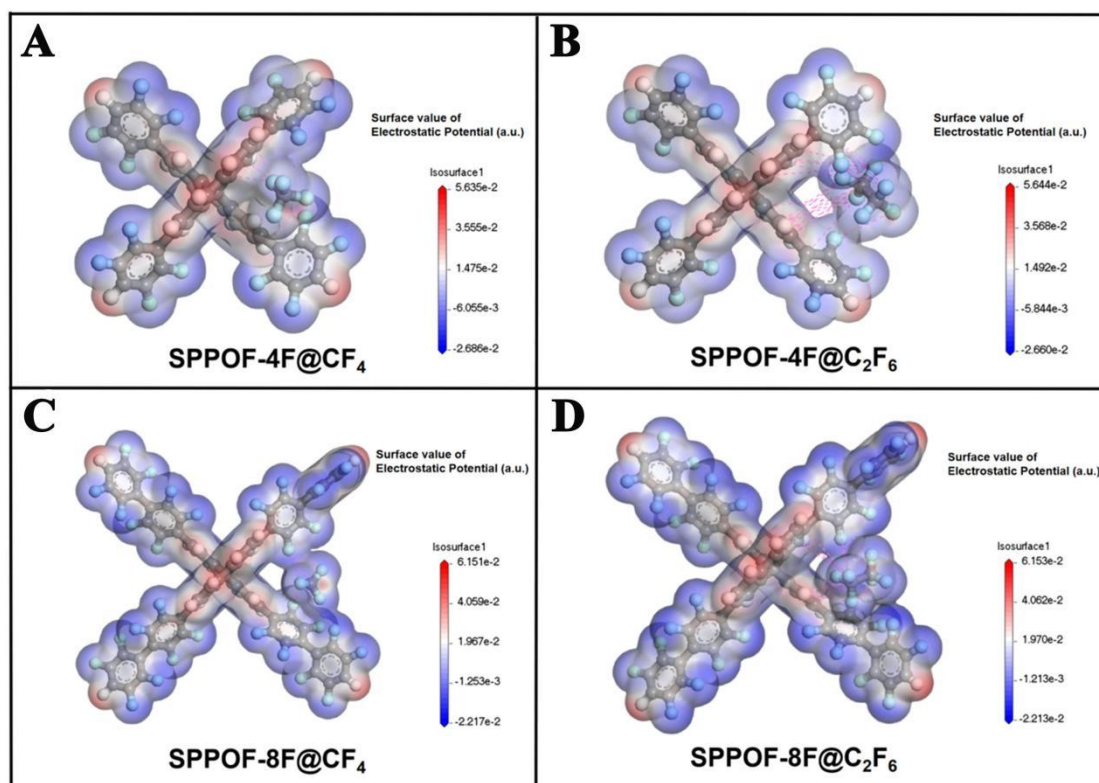
**Supplementary Figure 4.** TGA curves of (A) SPPOF-4F and (B) SPPOF-8F under air atmosphere.



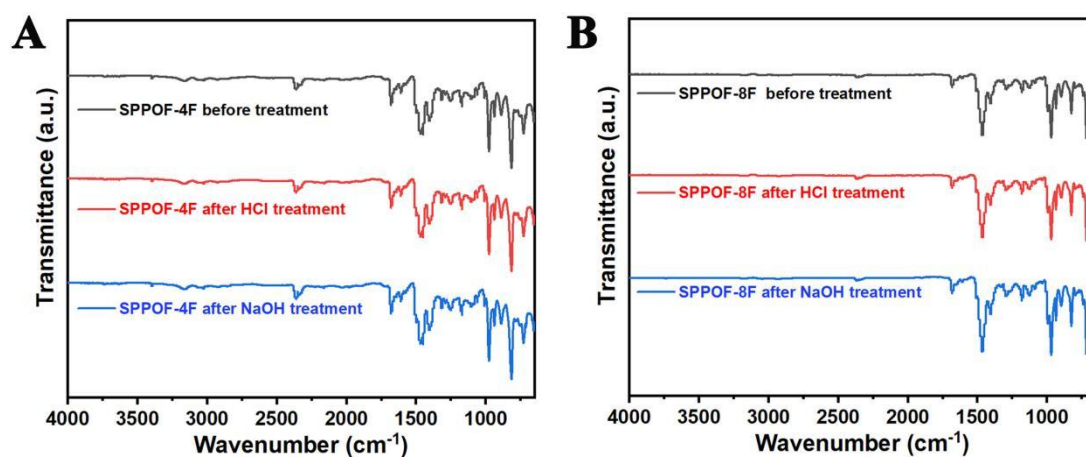
**Supplementary Figure 5.** Optical images of Contact Angle measurement for deionized water on the surfaces of (A) SPPOF-4F and (B) SPPOF-8F.



**Supplementary Figure 6.** Elemental distribution mapping images: (A) C for SPPOF-4F; (B) F for SPPOF-4F; (C) C for SPPOF-8F and (D) F for SPPOF-8F.



**Supplementary Figure 7.** Simulated model and the electrostatic potential maps of (A) SPPOF-4F@CF<sub>4</sub>, (B) SPPOF-4F@C<sub>2</sub>F<sub>6</sub>, (C) SPPOF-8F@CF<sub>4</sub> and (D) SPPOF-8F@C<sub>2</sub>F<sub>6</sub>.



**Supplementary Figure 8.** Comparison of FT-IR spectra of POFs before and after HCl or NaOH treatment: (A) SPPOF-4F and (B) SPPOF-8F.

## Supplementary Tables

**Supplementary Table 1. C<sub>2</sub>F<sub>6</sub>/CF<sub>4</sub> adsorption compared with literature**

<b>Adsorbents</b>	<b>Q<sub>st</sub> for C<sub>2</sub>F<sub>6</sub> (kJ/mol)</b>	<b>Q<sub>st</sub> for CF<sub>4</sub> (kJ/mol)</b>	<b>C<sub>2</sub>F<sub>6</sub> uptake (mmol/g)</b>	<b>CF<sub>4</sub> uptake (mmol/g)</b>	<b>Ratio of uptake (C<sub>2</sub>F<sub>6</sub>/CF<sub>4</sub>)</b>	<b>Note</b>
SPPOF-4F (This work)	29.0	20.5	1.74	0.79	2.20	298 K, 1 bar
SPPOF-8F (This work)	28.5	11.4	1.83	0.71	2.58	298 K, 1 bar
CS400 <sup>[1]</sup>	34.52	25.15	0.49	0.63	0.78	300 K, 1 bar
CS1000 <sup>[1]</sup>	40.69	31.45	0.33	0.41	0.80	300 K, 1 bar
CS1000a <sup>[1]</sup>	31.17	21.86	5.4	5.3	1.02	300 K, 1 bar
Al-Fum <sup>[2]</sup>	20.3	23.4	3.3	2.1	1.57	298 K, 1 bar
MOF-303 <sup>[2]</sup>	24.7	33.5	2.51	1.53	1.64	298 K, 1 bar
MIL-160 <sup>[2]</sup>	17.1	17.7	1.48	0.94	1.57	298 K, 1 bar
Zeolite 13X (8-12 mesh) <sup>[3]</sup>	14.0	24.4	1.55	0.75	2.07	303 K, 1 bar
Activated Carbon (20-40 mesh) <sup>[3]</sup>	20.6	19.4	1.26	0.75	1.68	303 K, 1 bar
Activated Carbon (12-20 mesh) <sup>[3]</sup>	---	---	1.21	0.68	1.78	303 K, 1 bar
Silica gel (30-60 mesh) <sup>[3]</sup>	---	---	0.50	0.14	3.57	303 K, 1 bar
H-cage <sup>[4]</sup>	---	---	0.62	0.39	1.59	273 K, 1 bar
HF-cage <sup>[4]</sup>	---	---	0.44	0.38	1.16	273 K, 1 bar
F-cage <sup>[4]</sup>	30.9	29.2	1.78	0.93	1.91	273 K, 1 bar

Supplementary Table 2. Mulliken atomic charges distribution in C<sub>2</sub>F<sub>6</sub>

Mulliken atomic charges of C <sub>2</sub> F <sub>6</sub>				
Atomic number	Atom	Mulliken atomic charges (a.u.)		
		C <sub>2</sub> F <sub>6</sub> alone	C <sub>2</sub> F <sub>6</sub> in SPPOF-4F	C <sub>2</sub> F <sub>6</sub> in SPPOF-8F
	C1	0.8128	0.8276	0.8292
	C2	0.8129	0.8305	0.8333
	F1	-0.2709	-0.2727	-0.2777
	F2	-0.2708	-0.2795	-0.2785
	F3	-0.2711	-0.2786	-0.2824
	F4	-0.2709	-0.2712	-0.2751
	F5	-0.2708	-0.2773	-0.2798
	F6	-0.2712	-0.2804	-0.2736
<b>Total atomic charge</b>		<b>0</b>	<b>-0.0016</b>	<b>-0.0046</b>



## References

- [1] Peng X, Vicent-Luna, et al. Separation of CF<sub>4</sub>/N<sub>2</sub>, C<sub>2</sub>F<sub>6</sub>/N<sub>2</sub>, and SF<sub>6</sub>/N<sub>2</sub> mixtures in amorphous activated carbons using molecular simulations. *ACS Appl Mater Interfaces* 2020;12:20044-55. <https://doi.org/10.1021/acsami.0c01043>
- [2] Zhu J, Hu J, et al. Aluminum-based metal organic frameworks for greenhouse gases CF<sub>4</sub> and C<sub>2</sub>F<sub>6</sub> capture with excellent capacity and selectivity. *Sep Purif Technol* 2024;331:125614. <https://doi.org/10.1016/j.seppur.2023.125614>
- [3] Ahn N, Kang S, et al. Adsorption isotherms of tetrafluoromethane and hexafluoroethane on various adsorbents. *J Chem Eng Data* 2006;51:451-6. <https://doi.org/10.1021/je0503756>
- [4] Tian K, Elbert, S.M, Hu XY, et al. Highly selective adsorption of perfluorinated greenhouse gases by porous organic cages, *Adv Mater* 2022;34:2202290. <https://doi.org/10.1002/adma.202202290>